

Properties and Application of Electrochemically Anodized Ta₂O₅ Nanostructures: A Review

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ABSTRACT

Anodization is one of the simplest synthesis methods that can be employed to generate generous nanostructures metal oxide. Knowing the potential of anodized Tantalum Pentoxide (Ta₂O₅) to form the desired nanostructure which is capable to produce outstanding performance, the elements and factors influence properties enhancement were gathered. The effect of parameter variation during anodization such as electrolyte composition, voltage variation, and anodization duration was evaluated to inspect the nanostructure formed. The results generated the differences in properties of nanostructures. Characterization methods such as field emission scanning electron microscopy (FESEM), X-ray diffraction (XRD), Transmission electron microscopy (TEM), Atomic force microscopy (AFM), and X-ray electron spectroscopy (XPS) were used to investigate the structural, electrical, chemical, optical, biological and mechanical properties of anodized Ta₂O₅. The findings show that anodized Ta₂O₅ is reliable to be employed as gas sensor, photocatalysis, pH sensor, biomedical materials and humidity sensor due to its superior properties to produce high sensitivity device and increase the efficiency. In this review, a general overview of anodized Ta₂O₅ is presented which focuses on its fundamental properties, method of synthesis anodized Ta₂O₅, anodization

parameter optimization, and recent applications in chemical and biological area, along with a discussion on future research directions relevant to this material. It is hoped that by synthesising nanostructures via anodization, the sensing performance will improve and Ta₂O₅ will be proven as a trustworthy metal oxide material for sensing applications.

Keywords: *Ta₂O₅; Synthesis; Anodization; Nanostructures; Sensors*

Introduction

Nanostructured materials have been broadly studied and enhanced in technologies for the last few years. In particular, transition metal oxide-based aluminium [1], titanium [2], niobium [3], and tungsten [4] were among the materials that have been widely implemented in various electronics applications and advanced catalysts. To synchronize with the nanotechnology industry that rapidly grows these days, these metal oxide materials are commonly studied and fabricated as nanostructures materials and presented in many forms to be employed in sensing devices or as energy storage such as the study of Nb₂O₅ nanochannel as photoanodes [5], W-Ta₂O₅ nanowires in photocatalysis [6], and nanoporous alumina in humidity sensing [7] with nanoscale dimension.

At the nanoscale level, innovative technologies, referred to as nanotechnology, have been compromised, making it feasible to design, produce, characterize, and regulate structures, materials, devices, and systems. Innovative advancements have been made in nanotechnology, especially in the field of nanomaterials, as researchers are discovering structures with varied and distinct attributes that are drastically enhancing numerous aspects of society in the past decades. Nanoscale-materials exist at different dimension structures ranging from zero-dimensional (0D), one-dimensional (1D), two-dimensional (2D) and three-dimensional (3D) nanostructures. Primarily, nanoparticles and clusters are categorized as 0D when all three spatial dimensions are at nanoscales while nanotubes, nanowires, or nanorods are 1D. 2D nanostructures are acknowledged when one of the spatial dimensions are nanoscales such as the structures formed on thin-film or nanosheets. 3D structures are formed when structural elements in 0D, 1D, and 2D materials are in close contact and generate interfaces. Porous nanostructures or compact polycrystals with nanosized grains are known as 3D nanostructures [8].

For many applications, metal oxide nanostructures can be synthesized via different approaches to maximize the surface area of interaction while altering the chemical and physical properties. Synthesis of nanostructure metal oxides can be conducted either in a solid-state, liquid state, or vapour phase deposition such as ball milling [9], hydrothermal [10], sol-gel [11], anodization [12], electrodeposition [13], and physical and chemical vapour deposition [14].

Each synthesis method encounter drawback in their nanostructure's fabrication. For example, hydrothermal method generates a small aspect ratio of 1D nanostructures and requires a few steps of synthesis while sol-gel method is time-consuming with the multistep process. However, these drawbacks in synthesis methods can be encountered through the anodization method which offers the solution based on the weakness of other synthesis methods [15]. Anodization is a form of properties enhancement method that is applicable for metal oxide. It is best known as one of the low-cost synthesis methods [16]. During the process, anodization parameters are varied to investigate various anodization conditions towards metal morphology, structure, and physical and chemical properties. This method is also simpler, able to be manipulated, and structurally tuneable [17]. Development of nanostructures ranging from 0D to 3D can be performed by varying the anodization parameters such as electrolyte composition, anodization time, and anodization voltage. In comparison to sol-gel technique which takes a long time to generate nanostructures on metal oxide, anodization is capable to generate nanostructures with similar dimensions in less duration [15].

Among numerous synthesis methods that allow generating of nanostructures, anodization provides high surface area and ease of repetitive fabrication, which can be considered as an advantage to this method in manufacturing metal oxide nanostructures. Narrowing to the potential of anodized material, outstanding physical and chemical properties can be generated from various anodization parameters which later can provide a favourable surface for the medium of interaction. Copper oxide anodized at certain anodization parameters can exist in the form of nanotubular, nanoporous, nanoneedles, and nanowires, which have the potential to be employed as chemical sensor, corrosion protection, photochemical water splitting, and solar light absorption [15]. There is evidence regarding the use of anodized nanoporous titanium oxide in photocatalytic activity [18]. Meanwhile, nanotubular titanium oxide has the potential to be employed as a biosensor in detecting SARS-CoV-2 as the cause of Corona Virus Disease (COVID-19) pandemic [19] and as bone implant [20]. The nanoporous coating has been fabricated recently from anodized aluminium oxide with the aid of sulphuric acid [1]. The outstanding properties of anodized material have drawn some attention to be used in the electronic industry as well as in biomedical applications.

Enhancement of properties through synthesis method for a particular metal can be a value-added as it improves the function as sensor or battery. Tantalum oxide (Ta_2O_5) is one of the transition-metal oxides that receive significant interest among researchers. In the past recent years, many studies have been conducted concerning Ta_2O_5 performance. Considering the potential of Ta_2O_5 and the anodization method to enhance nanostructures fabrication, the studies related to anodize Ta_2O_5 are gathered. Synthesized Ta_2O_5 has shown excellent results in improvising the properties of metal oxide after

undergoing anodization, which is believed to provide better performance in any practical application or device fabrication for various applications with aided of nanotechnologies. This has influenced the option of Ta₂O₅ over other metal oxides to be implemented in a wide range of applications such as sensors, coating, and biomedical technologies.

In order to explore the potential of anodized Ta₂O₅, its physical and chemical properties are obtained through different characterization methods such as FESEM, TEM, XRD and XPS which can visualize its behaviour. This paper discusses the influence of parameter variation during Ta₂O₅ anodization to alter significant changes in its physical and chemical properties which later improves the Ta₂O₅ performance in sensing application. A brief comparison is discussed in terms of electrolyte composition, voltage variation, anodization duration, and annealing condition as the factors affecting nanostructures. The discussion is projected to further explore the unique properties of anodized Ta₂O₅ obtained through characterization methods. Anodized Ta₂O₅ nanostructures applications in various industries are gathered to observe the effectiveness of anodization compared to other synthesis methods.

Anodization Parameter Optimization

Anodization is regarded as one of the most practical and straightforward procedures in forming extremely porous and ordered nanostructured metal oxide films. To carry out the anodization experiment, an electrode such as a platinum gauze is applied at the cathode and the substrate is attached to the anode. An electrolyte is poured across both electrodes. An anodization voltage is delivered between two electrodes during the synthesis. This causes electrochemical processes to occur on the surface of the substrate, resulting in the formation of a nanostructured metal oxide. The electrolyte composition, applied voltage, and anodization period are all essential parameters that influence the formation of nanostructures and the shape of anodic metal oxide films. Aside from that, metal oxide nanostructures are influenced by the annealing temperature, which occurs after anodization. As-anodized Ta₂O₅ films are mostly amorphous and require post-annealing to be turned into highly crystalline and modify the stoichiometric Ta₂O₅.

Effect of electrolyte composition

Anodization technique is implemented by preparing electrolytes with the adequate composition of the solution before the metal oxide compound can be immersed. From previous studies, electrolyte solution can be composed by combining sulphuric acid (H₂SO₄) and hydrofluoric acid (HF) [21], H₂SO₄, HF and water (H₂O) [22], H₂SO₄, HF, and ammonium fluoride (NH₄F) [23], glycerol and NH₄F [18], H₂SO₄, glycerine, H₂O and NH₄F [24], ethylene glycol, NH₄F, H₂O and H₂SO₄ [25], and NH₄F itself [17]. The concentration

of each element in the electrolyte composition has a significant influence on the growth of nanostructures.

There is a significant reason for adding H_2SO_4 in the electrolyte. It is said that H_2SO_4 is capable of forming nanotubular structure rather than the coral-like which is believed to have better performance in sensing application [24-25]. Increasing the H_2SO_4 saturation in the electrolyte will increase the porosity of nanostructure formation by enhancing the rate of chemical dissolution [26]. Optimization of H_2SO_4 concentration in the glycerol-based electrolyte was conducted by Abu Talib et al. where the concentration of H_2SO_4 was varied by 0%, 0.5%, 0.75%, and 1% [27]. FESEM image indicates voids and tubes of nanostructures best formed with 0.5 % vol of H_2SO_4 . Higher concentration is considered to disrupt the formation of nanotubular while no nanotubular can be formed without the content of acid in the electrolyte. In some materials, different types of acid produce different morphology of nanostructures. For example, Mir et al. used three different electrolytes with different types of acids at the same concentration in the anodization of alumina [7]. The pore diameter achieved by anodization of alumina in an electrolyte containing 0.3 M sulphuric acid, oxalic acid, and phosphoric acid was 20 ± 1 nm, 34 ± 1 nm, and 136 ± 4 nm, respectively. This proves that the manipulation in electrolyte elements and concentration is capable to achieve a wide range of nanostructures size.

Fluorine anion can be obtained from sodium fluoride, fluoric acid or ammonium fluoride to be aided in the formation of nanostructures. This statement is supported by Sloppy's calculation which indicates that the dissolution rate of Ta_2O_5 is slow without the presence of fluorine ion in an electrolyte [28]. This is because the dissolution rate of fluorine ion is dependent on its concentration, thus a high concentration of fluorine ion will produce a better etching process towards pore walls, hence, forming nanostructures with a larger diameter [29]. However, as reported by Karuppusamy et al., excessive fluoride ions in the electrolyte may cause no production of nanostructures due to the permanent dissolution of metal oxide [30]. Previous works of literature have reported the presence of NH_4F at different concentrations of NH_4F . The anodization electrolyte is prepared by varying the NH_4F for 0.45 wt%, 1.0 wt%, 1.35 wt% and 3.0 wt% [27]. The findings have shown that 0.45 wt% and 1 wt% of NH_4F content in electrolyte composition is insufficient to produce nanostructures while a sample with 3 wt% indicates that the structure is excessively etched and damaged.

An electrolyte requires organic additives in the formation of nanostructures. The presence of glycerine or ethylene glycol has its competency in increasing the viscosity and conductivity of electrolytes [24]. The presence of glycerine can also change the viscosity of the electrolyte. Given that viscosity is a critical element affecting ion mobility and consequently current density, it is altered by adding varying amounts of glycerol to produce different oxide layer thickness, pore diameter, and

interpore distance. Similarly, the higher the polyethylene glycol (PEG) content, the higher the electrolyte's relative viscosity as it produces lower pore width and thickness of anodic aluminium oxide [31]. As a result, variations in the electrolyte's viscosity are assumed to be a factor affecting ion migration, or perhaps the ionic current through the barrier layer.

In the anodization of Ta₂O₅, electrolyte composition is varied to achieve optimized and most promising nanostructures. To explore the role of water in electrolyte composition, Wei et al. work added an adequate amount of water into glycerol and NH₄F electrolyte. Compared to the aqueous electrolyte, the presence of water during anodization assists in decaying rapid current densities from the initial value [29]. It is strong evidence that the formation of compact oxide is promoted by the presence of water and requires high fluoride concentration to achieve porous layers. Referring to Momeni et al. [23], the researcher anodized Ta₂O₅ into two different types of electrolytes which are made up from the composition of H₂SO₄, HF and ethylene glycol and HF, H₂SO₄, and water respectively. The same anodization voltage of 15 V was applied to both samples. SEM images showed that anodizing Ta₂O₅ for 10 min produces irregular cracks while further elongation in anodization time produces parallel grooves and channels on the film surface. However, the compact film formed without porosity proves that anodization requires the presence of water to produce nanoporous or nanotubes structures.

From the result obtained by manipulating electrolyte composition, it can be said that solution concentration plays an important role in defining the metal properties after anodization. Excellent nanotubular structures can be achieved with an electrolyte composed of ethylene glycol, H₂SO₄, NH₄F, and water with NH₄F as the etching agent for pore formation with the addition of water and H₂SO₄ to increase chemical dissolution rates [27], [32].

Effect of voltage variation

According to Momeni et al.'s [21] research, the morphology of anodized Ta₂O₅ varies depending on the electrolyte composition, voltage applied, pH, and anodizing time. The considerable effect of voltage on the structure of Ta₂O₅ with the change of the voltage applied toward anodizing metal has been discussed in this work. Because the formation of nanostructures is closely related to anodizing voltage, voltage variations are required during anodization to produce an ideal condition in the formation of oxide layer nanostructures.

As anodized, nanotube array films start forming on the surface of tantalum. As higher voltage is applied, the structure and formation of nanotube array films are extended to higher order and become parallel to each other [21]. Nanoporous diameter starts to enlarge as more voltage is applied. Therefore, the formation of the nanostructure oxide layer is highly dependent on anodizing conditions, whether in the form of compact oxide layers, nanoporous, nanotubes, or fishbone structures. High anodizing voltage can directly affect the dissolution degree due to intensive etching which increases

the elongation of nanotube arrays [24]. Meanwhile, a longer nanotube array may result in the detachment of the nanotubes from the tantalum sheet. Therefore, it is critical to manipulate the parameters during the synthesis of tantalum nanostructure to produce a suitable nanotube array.

Momeni et al. [21] prepared two anodizing conditions which were operated at 15 V and 250 V for 120 s and 20 s respectively. After being anodized in the same composition of HF, H₂SO₄, and water as electrolytes and subjected to the corresponding voltage, two distinct samples were tested for corrosion resistance. In this experiment, the formation of nanotube arrays was successfully obtained as the nanotubes order was parallel to each other and perpendicular to the substrate. As higher voltage is applied, oxide layers start to further extend on tantalum result in the formation of nanoporous with a diameter range of 50 nm to 120 nm.

Another anodization of Ta₂O₅ was further investigated by Baluk et. al. [24] regarding the influence of anodizing voltage towards the formation of nanotube array and its adhesion. Few samples were distinguished based on their voltage variation, annealing temperature, and cleaning and drying conditions. Samples were anodized in similar electrolytic and annealing temperature conditions with an anodizing voltage of 10, 15, and 20 V which increased external diameter and length as voltage value increased. However, the adhesion of the nanotube array was weak when 20 V was applied where the nanotubes tended to peel off from Ta foil.

In the formation of nano-dimpled structures by Ma et. al. [12], increasing voltage can increase the diameter of nano-dimpled. Voltage varied between 10 to 70 V which produced nano-dimpled with diameter between 38 to 190 nm. Lee et al. [33] found that a lower voltage of 2 V is only capable to form non-organized pore order with a smaller size. However, the higher voltage applied such as 50 V only forms pores at the outer-most layer. Meanwhile, 20 V applied voltage is capable to form highly ordered pore formation. This is an ideal condition to achieve the desired outcome in K₂HPO₄ and anhydrous glycerol electrolyte setting.

To investigate the effect of voltage potential towards pore diameter and thickness, voltage ranging from 10 to 50 V was applied during anodization using the same electrolyte composition. The pore diameter of the nanoporous structure grew from 13 to 40 nm which can be said as linearly dependent on the voltage applied from 10 to 50 V, at approximately 0.64 nm/V. However, voltage potential higher than 50 V was unable to generate regular pore growth. At this point, the breakdown event took place. Meanwhile, the growth of porous thickness was not linearly dependent on the voltage potential as it increased from about 5.9 μm at 10 V to 15.7 μm at 20 V. There was a sharp decrease following after 20 V. In this case, this pattern is influenced by partial disbonding of layer at this particular voltage potential [29].

Chen et al. [34] reported that at 10 V, the Ta₂O₅ oxide layer appears as nanoporous. After 1 hour of anodization, nanotubes grew with a length of 5.2

μm and a diameter of 40 nm. However, tube formation did not elongate properly. The process continued by increasing voltage potential to 30 V and 40 V. As a result, the top diameter, bottom diameter, and the length of nanotubes at 30 V and 40 V were 57 nm, 72 nm, 20 μm , and 74 nm, 104 nm, 24 μm , respectively [34].

To conclude, a variation of anodization voltage leaves a significant effect on the formation of nanostructures in terms of size, length, and diameter. It is agreed that optimum applied voltage between 20 V to 40 V can stimulate oxide layer growth rates with proper arrangement of pore order [33].

Effect of anodization time

The time of the anodization is one of the parameters that need to be considered to form nanotubular layers. The formation of the parallel channel nanotubular structure indicates higher anodization duration is required compared to the formation of irregular crack structure obtained by shorter duration [23]. As the anodization duration is higher, the formation of nanotube will become obvious. Anodized layer of Ta_2O_5 will dissolve properly to achieve nanotubular structure due to increasing dissolution rate [25]. In most cases, nanotubes grow uniformly when longer duration takes place during anodization.

As reported by Lee et al. [33], at 40 s of anodization time, 10 to 18 nm of pore diameter and 6 to 18 nm of the wall thickness were achieved. At 3 min of anodization time, pore diameter and wall thickness grew until 10–13 nm and 24–28 nm with 20–25 nm of barrier layer obtained. Technically, increasing the anodization time enhances its morphology properties in terms of pore length, pore diameter and thickness.

To further explore the effect of anodization duration towards nanotubes growth, two samples with the same electrolyte composition and a voltage potential at 5 min and 10 min anodization time were compared. The difference in time interval results in different nanotubes diameter, thickness and length. For 5 min of anodization time, the diameter, thickness and length were recorded as 19 nm, 10 nm, 1.27 nm while 21 nm, 10 nm, 3.31 nm were recorded for 10 min of anodization time, respectively. The result obtained shows a further increment especially in terms of diameter and length when the anodization time is prolonged [24].

Hence, it can be said that nanotubes will further grow by lengthening the anodization duration. Anodization process is capable to enhance the morphological properties of Ta_2O_5 including the pore diameter, thickness and nanotube length. This is because, increasing the anodization duration will increase the dissolution rate of metal oxide layer, thus enabling the metal oxide anodized layer to dissolve appropriately during the formation of nanotubes.

Effect of annealing condition

Annealing is a process that involves the application of heat treatment by applying specific heat to metal oxide after the anodization process. As-

anodized samples produce an amorphous structure, and annealing must be carried out to achieve crystalline structure.

From previous studies, Ta₂O₅ underwent annealing in temperature between 150 °C – 1000 °C [17], [32], [35], [36]. A lower temperature may result in an amorphous structure while higher temperature might be able to form a crystalline nanostructure [17]. The higher the annealing temperature, the better the nanostructure formation which can also enhance adhesion of nanotubular [32], [37].

When the Ta film is subjected to anodization, the compact oxide layer may slowly disappear, develop nanoporous and nanotubular structures. Anodization takes place in a constant potential of 15 V for 30 sec to 5 min. As a result, nanotubular formed is well-adhered, with a diameter formation at 40 nm and tube length at 1 μm. After annealing takes place, diameter and tube length increase to 45 nm and 2 μm, respectively.

Knowing the potential of Ta₂O₅ in forming nanostructure at a different phase, a study was carried out to transform the amorphous nanoporous into a hexagonal phase. As-anodized Ta₂O₅ generates an amorphous structure on its oxide layer. XRD pattern in Figure 1 shows that after annealing at 290 °C, the structure transforms into a hexagonal structure which decreases the pore size ranging from 23 to 49 nm. Minagar et al. [38] also found that the hydrophilic property and the surface energy of the hexagonal nanoporous layers of Ta₂O₅ increased after the annealing process.

Most of the previous studies required high annealing temperatures above 300 °C to be applied on Ta₂O₅ to transform the amorphous phase into crystalline. This condition might be due to its high melting point that requires high thermal treatment to achieve the crystalline phase. Depending on the observation, the result indicates that the influence of thermal treatment plays an important role in the nanostructures phase to enhance crystallinity and alter Ta₂O₅ physiochemical properties [39].

As the potential of anodization has been acknowledged, this method was implemented to other metal oxides as well such as TiO, Al₂O₃, ZnO, and Nb₂O₅. Table 1 summarizes the effect of the anodization parameter on the morphology of nanostructures. The variation of anodization parameters in terms of electrolyte composition, anodization voltage, anodization duration, and annealing condition alters the dimension of nanotubes length, thickness, and pore diameter. The trend of synthesizing metal oxides using anodization has been increasing in the recent 5 years which proves that this is a reliable method to create novel nanostructures. However, some gap has been determined between the studies. Most of the nanostructures are fabricated using current anodization parameters generating nanoporous, nanodimpled, and nanotubular structures. There is no evidence reported on other kinds of nanostructures such as nanochannels, nanorods, and nanowires composed of anodized Ta₂O₅.

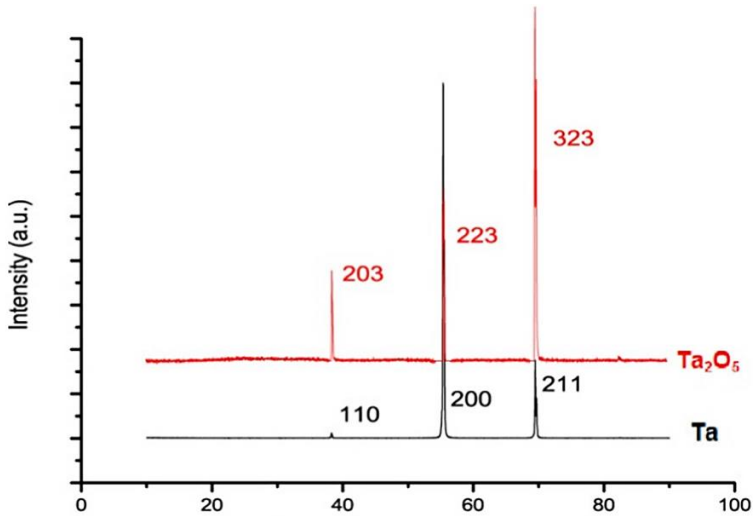


Figure 1: Comparison of XRD pattern between tantalum foil and Ta₂O₅ after annealing which shows the transformation from amorphous nanoporous into hexagonal phase. Permission open access from [38]

Anodization Mechanism in the Formation of Nanotubular Ta₂O₅

Formation of nanotubes from nanoporous structures

Anodization is an electrochemical method that is used to construct the naturally existing oxide layer on metal surfaces like Nb, Al, Ti, Ta, Cu, V, and Zr. It is often carried out in a two-electrode electrochemical cell, with the metal serving as the anode and platinum serving as the cathode. Oxygen ions generated at the electrolyte/anode interface move under the action of the electrical field, migrate within the metal attached to the anode and form an oxide surface layer when a potential difference between the electrodes is applied [40]. It is possible to manipulate electrochemical parameters, such as anodization duration and voltage, as well as electrolyte concentration, temperature, and pH, to change the characteristics of oxide layers and create novel surface morphologies like nanotubes, nanochannel, nanowires, and nanodimples on transition metal.

Table 1: Summary on nanotubes length, thickness and pore diameter of Ta₂O₅ formation according to electrolyte composition, variation, anodization duration and annealing condition

Morphology	Substrate	Electrolyte composition	Anodization voltage	Anodization time	Dimension
Nanotubes [18]	TiO film	Glycerol, H ₂ O, NH ₄ F	60 V	6 h	D: 90-100 nm T: 20-25 nm
Nanoporous [93]	Al ₂ O ₃ film	Oxalic acid	20 V, 40 V, 50 V	3 h, 1 h, 40 min	D: 31.8, 53.1, 71.0 nm
Nanoporous [94]	Fe ₂ O ₃ film	EG, NH ₄ F, H ₂ O	40 V	1 h	D: 59 ± 2 nm
Nanotubes [2]	Ti ₆ ⁴ plates	EG, NH ₄ F, H ₂ O	60 V	1 h	D: 120 nm
Nanoporous [40]	Ta film	H ₂ SO ₄ , HF	30 V	5 – 20 min	D: 25 – 65 nm
Nanoporous [17]	Ta ₂ O ₅ thin film	NH ₄ F, EG	15 V	2 h	D: 8–10 nm T: 3 nm
Nanoporous [78]	Ta ₂ O ₅ film	Millipore water, HF, H ₂ SO ₄	-	10 min	T: 45±2 nm;
Nanoparticles [95]	Zn film	HCO ₄	20 V	1 h	particles sizes: 10 – 20 nm
Nanotubes [29]	Ta film	Glycerol, NH ₄ F,	20 V	3 h	D: 17 nm
Nanotubes [29]	Ta film	Glycerol, NH ₄ F, H ₂ O (1 – 10 ml)	20 V	3 h	T: 39 nm – 138 nm
Nanotubes [23]	Ta film	H ₂ SO ₄ , HF, EG	15 V	10 – 20 min	T: 2 μm D: 55 nm
Coral-like [48]	Ta film	EG, H ₂ O, H ₃ PO ₄ , NH ₄ F	20 V	1 h	30 nm to 50 nm. L: 1.27 – 3.31 μm
Nanotubes [24]	Ta film	H ₂ SO ₄ , glycerine, H ₂ O, NH ₄ F	15 V	5 – 10 min	T: 10 nm D: 41 – 42 nm
Nanotubes [21]	Ta film	Hf, H ₂ SO ₄ , H ₂ O	15 V – 250 V	20 s – 120 s	D: 50 nm - 120 nm T: 6 nm – 25 nm
Nanoporous [67]	Ta film	K ₂ HPO ₄ , glycerol	-	40 s – 180 s	D: 10 nm- 28 nm

Nanotubes [29]	Ta film	Glycerol, NH ₄ F	10 V – 50 V	3 h	T: 5.9 μm 15.7 μm
Nanotubes [38]	Ta film	HF, H ₂ SO ₄ , H ₂ O	15 V- 250 V	-	D: 50 nm -120 nm
Nanotubes converted to nanodimpled [56]	Ta film	NH ₄ F, H ₂ O, diethylene glycol, H ₂ SO ₄	7 V – 70 V	-	L: 2–9 nm D: 25–60 nm

Ta₂O₅ nanotubular growth mechanism occurs in four stages that begin with the formation of the oxide layer by tantalum hydrolysis, following by pore deepening, initial nanotubular formation until nanotubular structures are perfectly formed, as summarized in Figure 2 [27].

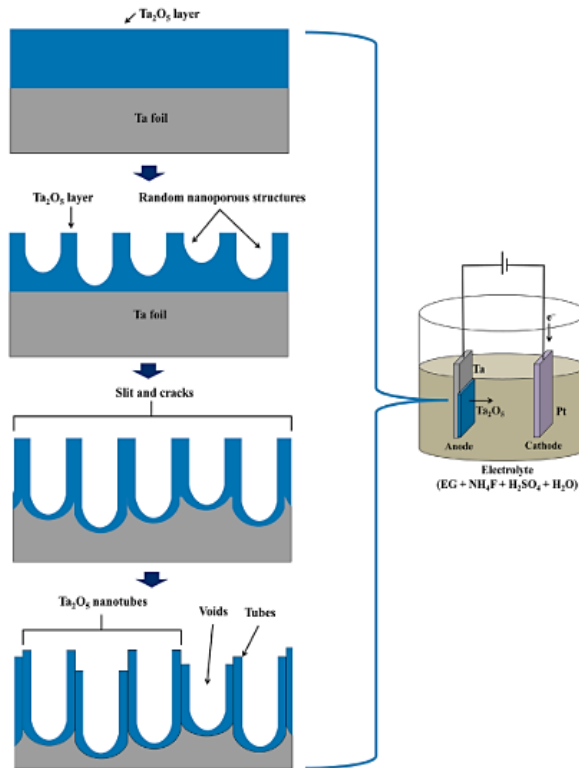


Figure 2: Illustration of Ta₂O₅ nanotubular growth mechanism during anodization. Permission open access from [27]

At the first stage, the oxide layer starts to form and due to its low electrical conductivity and electric field-enhanced dissolution of the oxide layer, this oxide layer significantly reduces the anodization current resulting in the formation of first pores. Development of the tantalum oxide layer and acid-induced dissolution processes leads to the development of nanostructured surface features in the anodization of tantalum. The nanotubular structure requires the presence of F^- ions, which cause chemical etching to generate porous and self-aligned Ta_2O_5 ordered architectures [41]. In contrast to other acids, sulfuric acid (H_2SO_4) is characterized by high thermodynamic stability, and it plays a key role in promoting the uniform development of nanotubes. F^- may be extracted from sodium fluoride or fluoric acid, which are among the most common chemical components. The $[TaF_7]^{2-}$ anions are an important component in the development of nanotubes. Ionic species such as F^- and O^{2-} migrates from the electrolyte to the metal-metal oxide interface while tantalum ions flow in the opposite direction under the applied potential to form Ta_2O_5 nanostructures. While optimizing anodization parameters to fabricate nanoporous surface features, it is important to acknowledge that higher etching rates are better achieved using a more concentrated electrolyte, resulting in the formation of nanotubular features. Meanwhile, lower etching rates are found when using a less concentrated electrolyte, resulting in a compact oxide layer. For the nanopores to form, it requires the electrolyte to be concentrated enough to promote the right concentration of F^- ions to partially dissolve the readily formed oxide layer. Hence, it is necessary to increase the electrolyte concentration so that the oxide layer dissolves in a way that would then form the nanopores on the surface.

Next, the dissolution of Ta_2O_5 begins, and pores develop following the separation of one pore from another, resulting in the establishment of distinct tubes that are vertically aligned with the substrate. The applied electric field promotes the production of field-assisted transport of Ta^{5+} and O^{2-} and the dissociation of O^{2-} from H_2O for which the chemical interaction leads to Ta_2O_5 being formed. Following the increment in anodization duration, fluoride ions are found deeper into the tantalum matrix and interact with tantalum, resulting in a higher anodization average pore size [40].

The formation mechanism of nanopores during anodization may be described by the $I-t$ curve in Figure 3 [40]. The occurrence of anodization produces a barrier layer on the tantalum surface that causes the current to raise in a short period. Following the development of the compact Ta_2O_5 layer in region a, a substantial reduction in current is noticed, since the current is unable to pass through the layer. Due to the dissolving reaction of fluoride ions, tiny pits appear on tantalum oxide in region b and transform into porous structures. Fluoride ions dissolve in the barrier oxide that causes the current going through the electrochemical cell to rise. The maximum current is attained at the end of region c. A steady drop in current occurs in region d as the current diffusion rate is limited due to the oxide formation and dissolution processes in the pores.

The pH value of electrolyte solution after anodization is slightly lower compared to before anodization takes place. This is because if the anodizing is kept on for a longer period, it could promote acid dissolution, which will then cause the adjoining pores to merge [42].

The length of NTs is highly dependent on the voltages, temperatures, and periods applied during anodization. To obtain a high etch rate and improve the dissolving rate of oxide layers, it is advisable to use a higher voltage during the anodic oxidation. However, a too high voltage supply may cause the nanotubes to be extracted from a tantalum sheet if they are longer than the length at which they can be pulled off. Thus, it can be assumed that applying the appropriate condition of anodic oxidation to a tantalum sheet results in a synthesizing of Ta₂O₅ nanotubes that adhere well to a substrate [24].

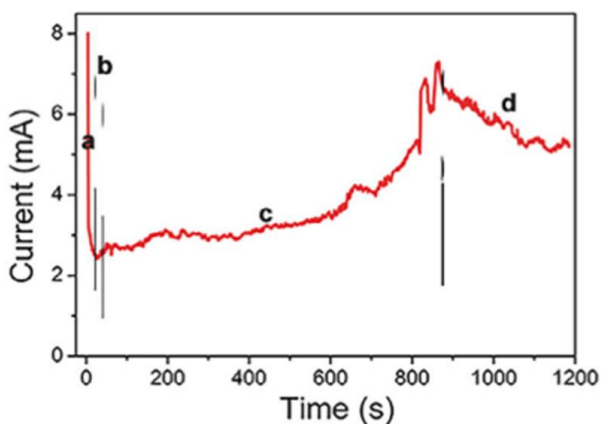
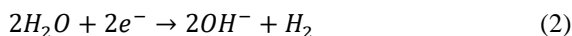
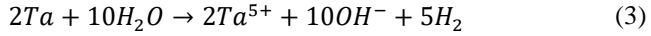


Figure 3: Current-time graph obtained during anodization of tantalum to reveal the current behaviour during the formation of nanoporous tantalum oxide layers up to 20 min anodization time. Reprinted from [40] 2020 with permission from Wiley Periodicals, Inc.

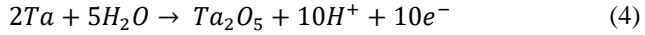
Chemical reaction during anodization

When neutral Ta is anodized in acidic aqueous electrolytes, an associated electric field is effectively capable of inducing the oxidation of neutral Ta with H₂O to facilitate the anodic transition of neutral Ta into Ta₂O₅ [43]. The half-cell reactions at anode and cathode and net reactions of the reduction process occur as in Equations (1), (2), and (3), respectively:

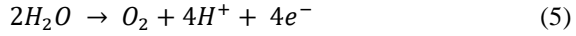




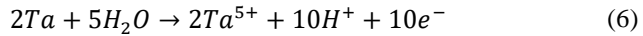
An oxide layer is built up on the metal by Ta hydrolysis at the electrolyte/metal interface, as described in Equation (4). Oxide barriers generated in anodic oxide can be fairly dense. The electrolyte may either dissolve or partially dissolve the anodic oxide. To produce a steady state of oxide formation and dissolution, an anodization parameter should be used to control the rate of oxide production and dissolution where porous metal oxide will grow under these conditions [44].



As seen in Equation (4), H^+ rises during hydrolysis, and F^- ions then travel to the H^+ sites to establish electroneutrality and fight for the O_2 sites in the oxide. To generate a tantalum oxide passivation layer on Ta foil, tantalum hydroxide is produced in tantalum hydrolysis. At a steady voltage, the oxide layer is created in the electrolyte. Acidification of the electrolyte could result in the development of large amounts of gaseous compounds because of water hydrolysis as shown in Equation (5).



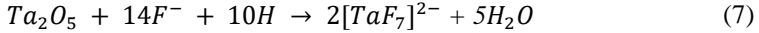
A regular arrangement of nanopores occurs, ranging from a cylinder form to the more generalized round, with a diameter of fewer than 100 nm. Equation (6) depicts the conversion of Ta to Ta^{5+} through oxidation.



Cation vacancies located at the metal /oxide contact allow for the reaction in Equation (6), which enables the Ta^{5+} to hop to the accessible vacancies. Nanopores form on the oxide surface while this condition is occurring. The concentration of F^- in the pores grows fast as H^+ ions are produced at the bottom of the hole where the Ta comes out of the metal and dissolves in the solution. Increasing in F^- exposure leads to the breaking down of porous walls until the neighbouring pores completely vanish. The size of pores also increases as the surface integration occurs [27]. At the same time, the increased electric field-assisted dissolution further enhances the process, therefore developing the voids. By stimulating the electric field, increased dissolution helped by the electric field is used to develop the pore, therefore developing the voids. The applied electric field promotes polarisation and metal oxide disintegration, as a result of which the Ta-O bond loses strength. This Ta is releasing ions which form acid in the pore bottom and affect pH profile developing in the pore development direction. Dissolution occurs

towards the bottom of the pore rather than at the top, resulting in a longer nanotubular.

Lastly, TaF_7^{2-} is formed from Ta_2O_5 by the local dissolution of H^+ , F^- , and O_2 to a critical level. Chemical dissolution influences the pore area grows as the Ta_2O_5 dissolving rate increases in fluoride-containing fluids, as stated in Equation (7). Negatively charged cation vacancies are formed in Ta after it has dissolved, and they migrate to the metal/oxide interface because of a potential gradient that forms across the oxide.



The oxide layer that forms when anodization occurs in an acidic solution is rather compact. Due to these limitations, self-organizing nanoporous oxide layers of these metals cannot be produced. Fluoride anions are necessary for the creation of self-organizing nanostructured coatings to solve this problem. Fluoride is unique in that it creates water-soluble metal-fluoride complexes, preventing the growth of metal-oxide layer build-up at the tube bottom. F^- ions, which have a greater migration rate across the oxide lattice, and have a lower ionic concentration than O^{2-} ions. This is what initiates the nanostructured coating growth since it induces the creation of a fluoride-rich layer at the metal-oxide contact. Figure 4 describes the process of the anodizing method of nanotube arrays development mechanisms of electrochemical anodization and the development of anodization electrolyte composition by generation [45].

Due to rapid oxide dissolution, the first generation of nanotube synthesis faced several challenges. Hence, HF was swapped with KF or NaF to improve the pH and increase the nanotube length in the second generation. With the third generation of synthesis, the length of the NTs was increased in non-aqueous electrolytes, such as ethylene glycol (EG), dimethyl sulfoxide (DMSO), formamide (FA), dimethylformamide (DMF), and N-methylformamide (NMF) with HF, NH_4F or KF mixed in. Currently, the fourth generation of nanotube synthesis has been found, which employs non-fluoride electrolytes and used in the anodizing of titanium [45].

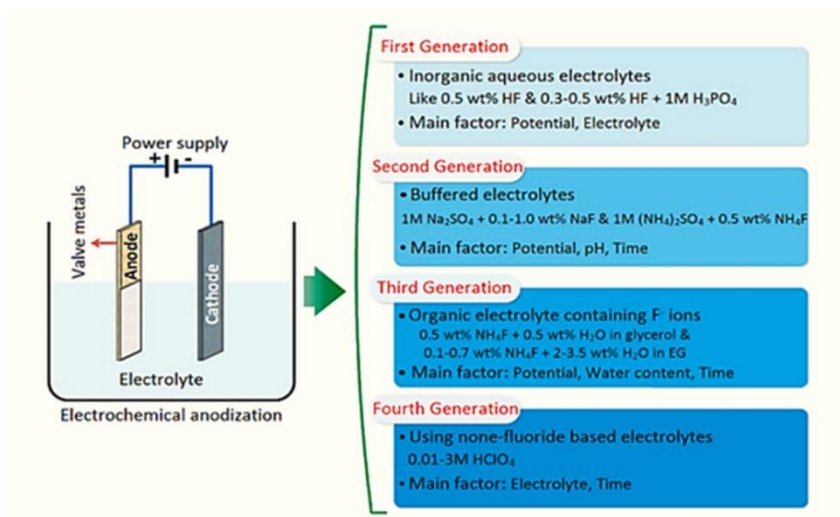


Figure 4: Different generation of TiO₂ nanotubular fabricated sing anodization synthesis method. Reprinted from [45] Copyright 2020 with permission from Elsevier

Anodization Parameter Optimization Properties of Ta₂O₅

Structural properties

Nanostructures of Ta₂O₅ are present in various structural features and shapes such as nanoporous, nanodimpled, nanotubes, and coral-like. The measurement of Ta₂O₅ morphology can be accessed through SEM, FESEM, or AFM. These analysis tools visualize the morphology and topography of nanostructures such as pore diameter, thickness of oxide layer, pore length, pore wall, surface roughness, and cross-sectional thickness.

Synthesis method such as electrospinning was used to synthesize Ta₂O₅ nanofibers Ta₂O₅ with average length and diameter of 40 nm and 12 nm, respectively [46]. As in the case of Ta₂O₅ nanospheres formation with ultrathin hollow spherical shell in size ±5nm, a facile template method has formed Co₃O₄/ Ta₂O₅ hollow nanospheres with the diameter ranging from 100 nm to 300 nm [47]. There was evidence from the one-step hydrothermal method to construct Ta₂O₅ nanorods and W Ta₂O₅ nanowires. The length and diameter of these nanowires were reported to be 0.7-1.2 μm and less than 30 nm, respectively [6].

Similar to other synthesis methods, different nanostructures of Ta₂O₅ can be obtained through anodization. As discussed in the previous section, electrolyte composition is one of the anodization parameters that contribute to

a different nanostructures formation to alter the morphologies of Ta₂O₅. For instance, the use of phosphoric acid (H₃PO₄) is said capable to form coral-like of Ta₂O₅ [14], [25]. In a study conducted by Yu et. al. [48], anodization electrolyte was composed of a combination of Ethylene Glycol, water, H₃PO₄, and NH₄F that resulted in the distribution of uniform nano-scaled pores on Ta₂O₅ film in an assembled morphology of coral-like structure. Nanoporous Ta₂O₅ with pore diameter ranging from 5 to 35 nm and oxide layer thickness of about 400 nm was anodized in an electrolyte containing HF or H₂SO₄. The formation of pores was randomly located with narrow size distribution. The distribution of pores, size and thickness can be manipulated by varying the electrolyte concentration, anodization time and voltage applied during the anodization reaction [14]. Nanotubes can be grown from the increment in voltage potential or duration during anodization [34]. Replacement of HF to NH₄F is a strong factor for the growth of nanotubular from nanoporous structures [45]. Minagar et al. [38] obtained the top view, bottom view, and cross-sectional view of nanotubular Ta₂O₅ with the diameter ranging from 35 to 65 nm and nanostructure thickness of $1.17 \pm 0.05 \mu\text{m}$. Wen et al. [49] reported that the morphology of Ta₂O₅ nanotubes produced with different current densities and anodizing times differs, which is primarily reflected in the nanotubes' length and diameter. The length of the nanotubes increases as the anodizing current density and time increase. The results show that Ta₂O₅ nanotubes of 18 μm length can be grown in less than 2 minutes at a current density of 125 mA cm⁻². Under constant current anodizing conditions in the H₂SO₄-HF electrolyte system, Ta₂O₅ nanotubes grow at a faster rate than under constant voltage anodizing conditions. These findings could pave the way for large-scale production of Ta₂O₅ nanotubes in the future. Figure 5 displays coral-like, nanoporous, and nanotubes structures as the formation of the potential nanostructures during anodization synthesis of Ta₂O₅.

Initially, the nanoporous structure started to form compact oxide layer once anodization began. Prolong anodization transformed the nanoporous structure to nanotubular. Other than structure formation, anodization time gives a different effect on Ta₂O₅ nanotubular surface roughness that is measurable through AFM which displays surface topography of the oxide film based on colour intensity to indicate roughness. Anodizing at 0.5 h produced surface roughness of 38.90 nm while anodizing at 1.0 h and 1.5 h produced surface roughness of 63.93 nm and 95.94 nm, respectively. It is said that higher surface roughness is capable to adsorb a higher number of oxygen molecules which results in rapid photoconductivity activities and enhancement in responsivity [25].

Another evidence has reported on topography properties of anodized Ta₂O₅ film treated with and without n-octadecyltrichlorosilane (OTS) which has been analysed through AFM. Result acquired has indicated that Ta₂O₅ film without OTS has 2.5 nm surface roughness while Ta₂O₅ film with OTS has 1.3 nm surface roughness. The anodized oxide layer is grown in a homogenous,

uniform manner and appears to be smoother that facilitates the low threshold voltage organic thin-film transistors with low surface roughness [50]. The literature has shown that anodization is capable to generate a wide range of Ta₂O₅ surface roughness between 1 to 100 nm that can be utilized in the various application according to its compatibility.

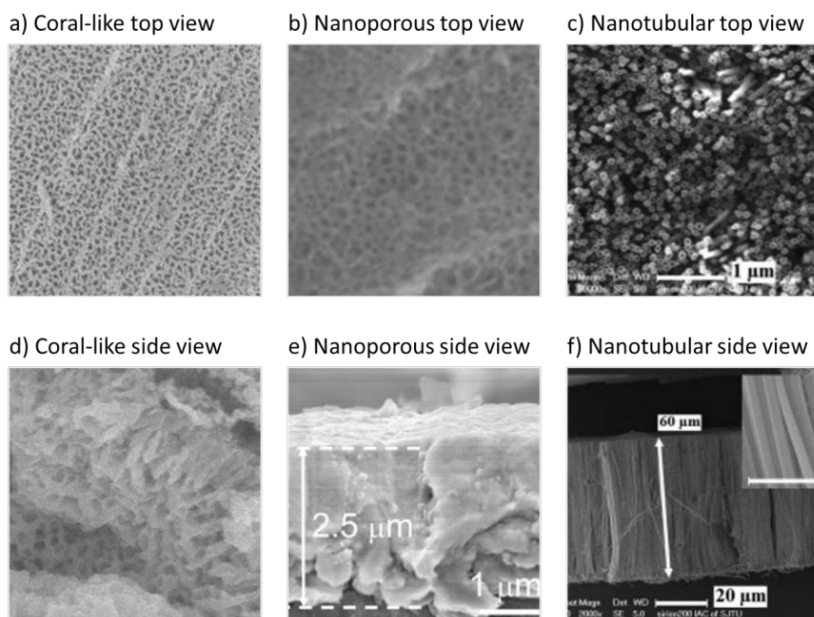


Figure 5: FESEM and SEM images of top view and side view Ta₂O₅ with three different types of nanostructures grown from different electrolyte composition by comparing (a) coral-like top view [48], (b) nanoporous top view [17], (c) nanotubular top view [34], (d) coral-like side view [48], (e) nanoporous side view [17] and (f) nanotubular top view [34]. (a and d) Permission open access from [48]. (b and e) Reprinted from [17], Copyright 2018, with permission from Elsevier. (c and f) Reprinted from [34] Copyright 2017, with permission from Elsevier

From this evidence, we can see the importance to manipulate anodization parameters to create a novel structure of metal oxide with distinct dimensions favourable for respective applications. This has justified that anodization is a good alternative for the synthesis of nanostructures into a different form such as nanotubes, nanoporous, nanodimpled and coral-like as the parameters are feasible to be manipulated. Based on previous studies, this simple method has successfully generated pore dimension at the nanoscale that can be employed accordingly to different applications.

Chemical properties

After anodizing in a sulphuric acid-containing electrolyte, the nanostructure of Ta₂O₅ contains impurity sulphur and excess oxygen. The presence of these impurities might probably cause the structure to become more stable and enhance its stability properties [51]. Once anodization takes place, the formation of a compact oxide layer is presented on the tantalum surface. Tantalum ions (Ta⁵⁺) released from tantalum metal become soluble once they reach the oxide layer and electrolyte, thus forming [Ta₇]²⁻. Hence, a pore is formed between the oxide layer interface and electrolyte. The pore formation further develops during the process as anodization assists to trap the necessary ions. Growth rates of anodized nanostructures such as nanoporous and nanotubes are influenced by the rate of diffusion of fluorine ions and metal complex solubility [38].

In comparison to other anodized transition metals such as titanium, tantalum has a higher rate of oxidization. Formation of tantalum nanoporous mechanism has a close relation with [Ta₇]²⁻ molecular dimension, Ta₂O₅ formation kinetics, and formation enthalpy. Low electrode potential is required to oxidize tantalum compared to titanium. In terms of forming enthalpy, tantalum releases more calories than titanium. Hence, it is highly agreed that the rate of oxidization of tantalum and formation of an oxide layer on the substrate surface is faster than titanium where its standard formation enthalpies are -492.790 and -228.360 gram calories per mole, respectively [38].

Compared to anodized Nb₂O₅, Ta₂O₅ possesses higher lithium storage capacity due to its Ta-O bond that is less robust than Nb₂O₅ in Lithium batteries application [5]. The amorphous structure of Ta₂O₅ develops a narrower bandgap and conveys desirable electrochemical activity. Furthermore, oxidation of tantalum occurs at different oxidation states such as +5 (Ta₂O₅) and +4 (TaO₂). Oxidation of +5 (Ta₂O₅) can be said as the most stable state. Passivation layer formed on the tantalum surface when permitted to air exposure results in unparalleled corrosion resistance. This is why tantalum can strongly resist chemical attacks from strong acid hydrochloric acid or nitric acid. Hence, strong etching agents such as hydrofluoric acid, fuming sulfuric acid and potassium hydroxide are capable to etch Ta₂O₅ [52].

XPS analysis visualizes the elemental and chemical bonding state as the characterization of Ta₂O₅ [32]. This method can be used to prove the formation of an oxide layer after anodization which is indicated by the photoelectron intensity peak of the anodized region binding energy for both tantalum and oxygen atoms [49], [52]. The measurement was done by measuring atomic concentrations for Ta and O atoms [33]. Corresponding to the Ta 4f_{7/2} and Ta 4f_{5/2} orbitals, synthesis of Ta₂O₅ via one-step heating has obtained Ta 4f element approximately located at two peaks centered at 26.5 eV and 28.4 eV using XPS. Meanwhile, O element was located at 530.4 eV and 532.3 eV corresponding to 1s orbital [54]. Figure 6a shows the XPS spectra for elements in Ta₂O₅. A facile template method used to synthesize Co₃O₄/ Ta₂O₅

Crystal structure

Similar to other metal oxides, Ta₂O₅ has polymorphic transformation of about 10-15 which has been reported until this date and can be identified by XRD [37], [57]. As-anodized structures are primarily amorphous and require post-heat treatment to transform into a crystalline structure. β -Ta₂O₅ phase can exist at low temperature while phase transition to α -Ta₂O₅ requires a high temperature above 1000 °C [41], [58]. Crystal structures of metal oxides have commonly existed in the form of triclinic, monoclinic, orthorhombic, tetragonal, trigonal, hexagonal, and cubic. One of the impressive Ta₂O₅ crystal attributes formed via anodization is its potential to generate wide ranges of crystal structures using Ta foil and Ta₂O₅ powder as substrate. Monoclinic crystal structure is commonly found at low-temperature post-heat treatment [6] while hexagonal and orthorhombic can be achieved at a high temperature of above 550 °C [35], [53], [59], [60], [61]. For most cases, anodization possesses a slightly similar trend where annealing at 400 °C produces cubic structure while increasing the temperature up to 750 °C produces tetragonal structure and following temperature increment of above 800 °C can generate orthorhombic crystal structure as shown in Table 2 [42], [50], [54], [62].

Regardless of the anodization voltage, Ta₂O₅ nanotube arrays grow in an amorphous phase until the sample is being annealed [34], [38]. Annealing at high temperatures will result in the transformation of amorphous oxide into the crystalline phase. All diffraction peaks are identified, and the largest five peaks are determined to correspond to degrees of crystallinity, which indicates the crystallization behaviour of the sample [51]. Initially, annealing at 1058 K is found to be in the amorphous state. However, after annealing in 1128 K, the crystalline structure of orthorhombic β - Ta₂O₅ is able to form diffraction peaks identified at 22.9°, 28.4°, 36.8°, 46.7°, and 55.5° [51]. There is also evidence that annealing in 700 °C is able to form hexagonal phase with peaks appearing at 23°, 28.5°, 36.5°, 50.5°, and 55.5°. Figure 6b visualizes the Ta₂O₅ peak for orthorhombic structure. By increasing annealing temperature to 800 °C, the crystalline structure transforms from hexagonal to orthorhombic with diffraction peaks at 50.5° being split into 50° and 51° [64].

Another characterization method that can be used to characterize the crystal phase of Ta₂O₅ is TEM imaging. Variation of chemical phases including measurement of nanoparticles size, size distribution, grain size, and morphology with clear interfaces can be obtained. Total sample thickness, period, and sub-layer thicknesses are also possible to be measured through TEM [65]. There is a slight difference in image formation using TEM and SEM which can be seen in Figure 7 since the TEM images are created using transmitted electrons whereby the electrons are passed through the sample to create an image. The microstructure layer of Ta₂O₅ is characterized as amorphous with a gap value of 2000 Å (< 100 Å) [66]. Different chemical interfaces can be observed through TEM analysis where the interplanar distance of pure Ta is 2.33 Å; thus, it identifies crystal plane 111 of α -Ta for

all crystallinity [65]. TEM is able to provide porosity of selected area diffraction pattern in which as-anodized Ta₂O₅ porous structure can obtain porosity of 26% which indicates its amorphous nanostructures [33]. Crystal plane of 020 with fringe interspace of 5.1 Å is identified as orthorhombic [16].

Table 2: Crystal structure according to their post-heat treatment temperature based on substrate and synthesis method

Substrate	Synthesis method	Crystal structure	Post-heat treatment temperature (°C)	Ref
Ta ₂ O ₅ and S-doped Ta ₂ O ₅ nanocomposites	Sol-gel method	Orthorhombic	700	[61]
Ta ₂ O ₅ powder	Facile one-step heating	Orthorhombic	550	[54]
TaCl ₅ powder	Ultrasound-assisted	Orthorhombic	800	[60]
Ta ₂ O ₅ doped tungsten	Hydrothermal	Monoclinic	200	[6]
Ta ₂ O ₅ powder	Hydrothermal	Hexagonal	550	[62]
Ta ₂ O ₅ deposited on silicon	Ion beam sputtering deposition	Hexagonal	800	[35]
Ta ₂ O ₅ deposited on silicon	Ion beam sputtering deposition	Orthorhombic	1000	[35]
Ta ₂ O ₅	Template-assisted	Orthorhombic	700	[96]
Ta foil	Anodization	Orthorhombic	800	[51]
Ta foil	Anodization	Hexagonal	290	[38]
Ta foil	Anodization	Cubic	400	[42] [55]
Ta foil	Anodization	Tetragonal	750	[42] [55]
Ta ₂ O ₅ powder	Anodization	Cubic	400	[63]
Ta ₂ O ₅ powder	Anodization	Orthorhombic	800	[63]

Biological properties

Ta₂O₅ possesses outstanding thermal and chemical stability and its biocompatible properties are promising. Thus, its capability to be used as a biosensor and biomaterial is undoubted. Another interesting property exhibited by Ta₂O₅ is low toxicity due to a high oxidation state. Ta₂O₅ is highly inert and

biocompatible which makes it widely used in biomedical applications as biomaterial as well as resistance to acid corrosion [11], [14]. Until today, there has been no evidence regarding its cytotoxicity that may harm osteogenic cells and mesenchymal stem cells. In biomedical application, a natural protective coating is produced by forming oxide layers on the surface of tantalum during anodization which relatively provides extra protection for cells. Although the amorphous state of nanoporous is unable to generate nanotubular, this state will form linkages or platforms for allowing it to combine with certain drug during the drug loading process [64]. Ta₂O₅ has also been widely used as implant material based on its biocompatible properties [32], [66], [67].

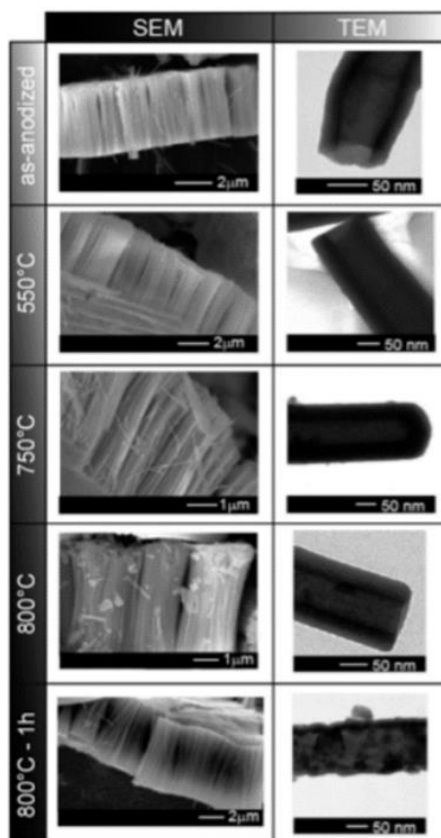


Figure 7: Comparison on image formation of Ta₂O₅ nanostructures, as anodized and annealing at different temperatures obtained using SEM and TEM. Reprinted with permission from [39] Copyright 2012 American Chemical Society

Compared to Nb_2O_5 , TiO_2 , and ZrO_2 , Ta_2O_5 has the highest bioactivity of biocompatible nanoporous and nanotubular oxide metals in simulated body fluid at 37°C measured after 3 weeks. This performance is indicated by the atomic ratio of calcium to phosphate due to absorption of Ca^{2+} ions onto the hydrolysed nanoporous and nanotubular oxide surface following the adsorption of existing phosphate groups to the positively charged surface inside the simulated body fluid; thus, forming calcium phosphate. Deposition of hydroxyapatite (HA) is observed on as-formed and annealed Ta_2O_5 film as shown in Figure 8.

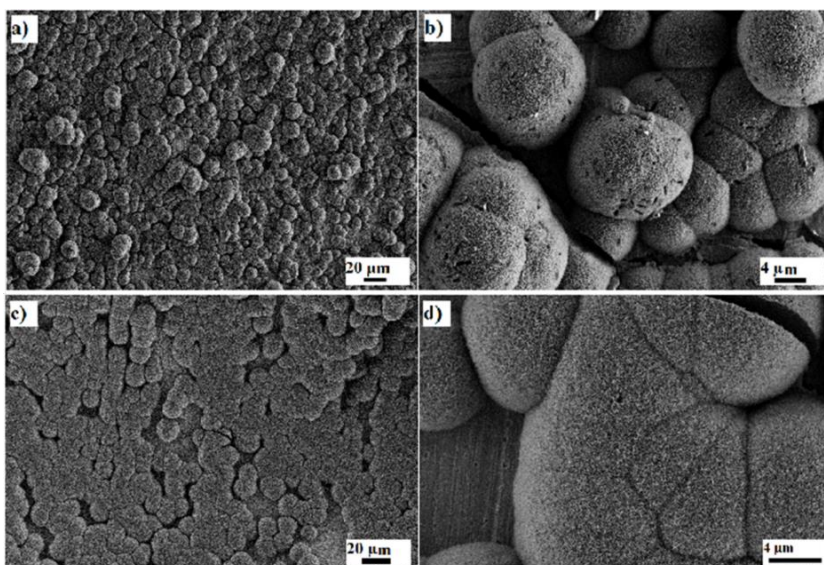


Figure 8: SEM images of (a) low magnification and (b) high magnification Ta_2O_5 deposited with Hydroxyapatite (HA) on (c) formed and (d) annealed film. Permission open access from [38]

Referring to Figure 9, the performance of Ta_2O_5 is indicated by the presence of a mixture of calcium and phosphate in compositions HA of bone composition [38]. It was found that biological properties possessed by anodized nanoporous tantalum oxide promoted fibroblast adhesion and proliferation, as well as the growth of fibroblasts on the anodized surfaces. Strong cellular adhesion on nanoporous tantalum with the higher surface interaction was recorded by tantalum nanoporous with 25 nm pore size [40]. Anodization of tantalum alters its surface topography and enhances the hydrophilicity of the porous tantalum [12]. On the other hand, nanodimpled Ta_2O_5 produced via anodization exhibits higher hydrophobic properties than

pure tantalum. Nanodimpled with a diameter of 90 nm has the highest cellular activity of adhesion towards mouse cranial anterior bone cells. Hence, the properties of anodized Ta₂O₅ can be manipulated and fabricated according to its favourable condition for application.

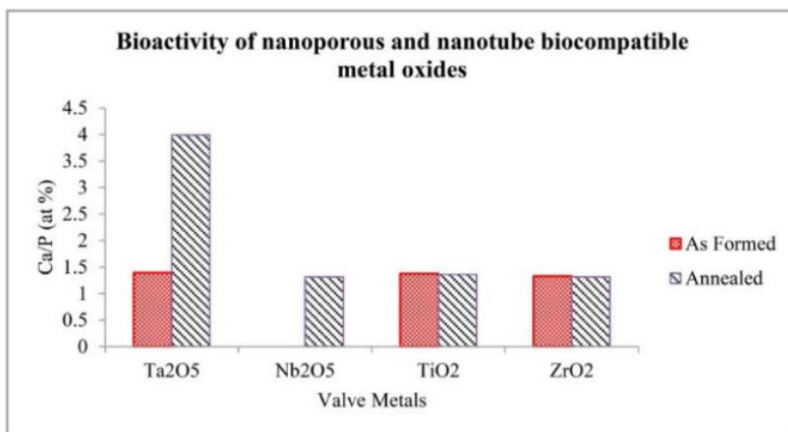


Figure 9: Atomic ratio of calcium to phosphate indicates by hydroxyapatite (HA) deposition to measure bioactivity performance of metal oxide after 3 weeks in modified simulated body fluid (m-SBF) between as-formed and annealed Ta₂O₅, Nb₂O₅, TiO₂ and ZrO₂. Permission open access from [38]

Optical properties

The reason why Ta₂O₅ is one of the best candidates for the optic-based sensor is that it exhibits high refractive index material which might enhance the sensitivity of the sensor [21], [57], [68], [69]. The presence of defects level and oxygen vacancy on its surface has aided in productive interaction by quickly adsorbing analyte molecules in the sensing surface [71].

To demonstrate the optical energy band gap of Ta₂O₅ nanoparticles prepared by ultrasound-assisted synthesis, UV–VIS diffused reflectance spectroscopy was used to investigate its optical properties. The study projected the bandgap for Ta₂O₅ nanoparticles as 3.12 eV which was slightly lower than the bulk Ta₂O₅ (3.8 eV) [60]. The green light emission of this Ta₂O₅ nanoparticles appears to be at its most intense level at a wavelength of 535 nm with a particle size of 20 nm. Since these nanoparticles have excellent luminescent properties, they can be used in optical devices and high-quality monochromatic lasers. Electronic band structure of β -Ta₂O₅, δ -Ta₂O₅ and orthorhombic with lattice parameter $a = 7.9 \text{ \AA}$, and $b = c = 3.75 \text{ \AA}$ were applied [72]. Calculated band gap produced 2.45 eV, 2.92 eV and 3.7 eV respectively. As Ta₂O₅ can exist in many crystals form in which each structure suits the certain condition of application, the existing of various Ta₂O₅ form produces

different band gap which allows optical properties control according to its functional suitability. Ta₂O₅ has a minimum absorption at ultraviolet light, which is known as the adsorption edge around 320 nm. This occurs due to the intrinsic band gap of Ta₂O₅ being around 3.9 eV [60], [72].

As the amount of Ta doped on the TiO₂ layer increases, the overall photocurrent density decreases. The gradual enhancement of the optical band gap in combination with decreasing photocurrent is consistent with the Ta concentration (from 3.31 to 3.44 eV). In this application, an ultrathin Ta-doped TiO₂ layer was applied to exhibit photoelectrochemical activity. This layer responded with photoelectrochemical currents of about 10–100 mA/cm² upon exposure to UV light (wavelength 320–380 nm; intensity 2.5 mW/cm²) [74].

According to Nickel et. al., the composite refractive index of metal oxide can be enhanced either by increasing anodization voltage or with a certain mixture of anodization electrolyte composition to increase the pore diameters [75]. The addition of chemical additives in anodization might also agitate to increase the refractive index by decreasing the dissolution rate. From the study, it was found that anodization of aluminium-tantalum thin film in 0.4 M phosphoric acid at 4 V was able to generate a refractive index of 1.258 while additional of 0.1 M oxalic acid and 0.4 M oxalic acid into the electrolyte mixture is capable to increase the refractive index to 1.319 and 1.359 respectively. Chemical additive such as oxalic acid reduces chemical etching; hence, decreases the pore size which reduces the air content in the thin film. As a result, the refractive index of the porous aluminium oxide layer increases due to a decrease in air content within the oxide layer [75]. Modification of optical properties is possible through anodization as synthesizing the nanostructure will alter the electronic band structure.

Mechanical properties

Porous tantalum is highly advantageous due to its mechanical properties especially in biomedical applications compared to solid tantalum. It was found that the elastic modulus of porous tantalum was 1.3–10 GPa which was lower than solid tantalum and titanium alloy which were 186 GPa and 106–115 GPa respectively. Meanwhile, the elastic modulus of porous tantalum was nearly close to cortical (12–18 GPa) and cancellous (0.1–0.5 GPa) bone tissue. The stress shielding effect can be reduced as it possesses a low elastic modulus [14].

The porosity of porous tantalum was 75%–80% while the porosity of other porous materials was 75%. Higher porosity value regulates tissue permeability and enhances the exchange of nutrient inside scaffold which potentially leads to tissue growth [14]. The porous structure exerted by Ta₂O₅ has approximately similar mechanical properties of natural bone.

Compression tests carried out using a servo-hydraulic MTS (axial/torsion materials test system) machine found that Young's modulus of laser-processed Ta samples can be varied in the range of 2–20 GPa [68]. In

comparison to the actual human trabecular bone, porous tantalum fabricated through chemical vapour infiltration and deposition and powder metallurgy technique is capable to create approximate properties as shown in Figure 10.

Meanwhile, Ta₂O₅ was seen to improve in corrosion behaviour after annealing took place as reported by Sarraf et al. [32], where it was analysed by using nanoindentation device result in maximum hardness of ~ 7.5 GPa and elastic modulus of ~ 160 GPa. Compared to its bare substrate, the modification for corrosion resistance was almost 50% and 29% higher. In addition, the enhancement of Ta₂O₅ durability properties can be done through the manipulation anodization duration [24].

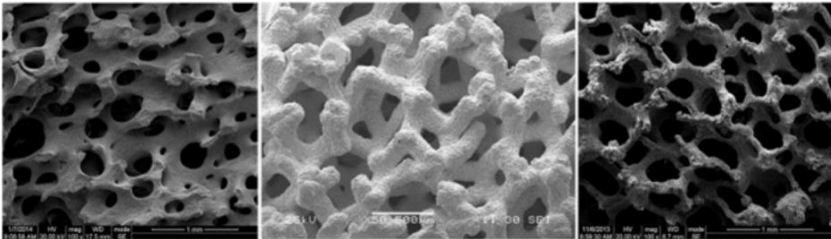


Figure 10: Image of Scanning electron micrograph (SEM) displaying actual human trabecular bone (left), porous tantalum fabricated through chemical vapour infiltration and deposition (CVI/CVP) (middle) and powder metallurgy (PM) (right) technique. Reprinted from [14] Copyright 2015, with permission from Elsevier

Application of anodized Ta₂O₅

In this section, the use of anodized Ta₂O₅ in chemical and biological applications is presented. Based on the literature conducted, Ta₂O₅ use is relatively small; thus, many material possessions can be explored and various novelties can be obtained with a more comprehensive and detailed study. In addition, the application of Ta₂O₅ synthesized using other methods is also discussed and compared to its performance.

Gas sensors

Ta₂O₅ has been found to exhibit assorted functions in sensing applications due to its extensive sensing properties. In relation to metal oxide gas sensors, it has been demonstrated that the use of nanoheterostructures can enhance their performance as the doping mechanism has improved the oxygen vacancy available on the structure [76]. While doping increases the concentration of oxygen vacancies, this has been shown to enhance chemisorbed oxygen concentrations.

A study focused on a capacitive hydrogen sensor made of an ultra-permeable Ta₂O₅ dielectric film on a SiC substrate tested with the range of

hydrogen concentrations from 0 to 2000 ppm in room temperature to 500°C. After 60 seconds of exposure to 2000 ppm hydrogen gas at 500°C, the capacitance increased from 5.0 to 6.5 nF. The proposed hydrogen sensor using a thin Ta₂O₅ dielectric film showed good sensitivity and adaptability at high temperatures. Pd-capped tantalum is excellent optical hydrogen sensing material [77]. The sample exhibits a much-defined optical response to hydrogen, with excellent resolution. There is no hysteresis observed and the optical response of the materials to hydrogen is stable over time.

Another test study conducted by Imbault et al. investigated the efficiency of Ta₂O₅ nanoporous as a permeable oxide membrane for gas sensing [78]. The fabrication of the sensor includes the placement of the Ta₂O₅ membrane on the top layer of platinum electrodes through the implementation of the Lift-Off-Float-On technique. The response of Ta₂O₅ towards carbon monoxide, hydrogen, ethanol, and ammonia was recorded and its efficiency was analysed in an active sampling of diluted analytes. A sensing membrane of 1 cm in diameter was capable to sample the airflow of a few hundred milliliters per second to obtain the presence of analyte molecules. It is reported that the use of Ta₂O₅ to construct a nanoporous membrane for gas sensing devices is due to its properties that exert high redox chemical activity at elevated temperatures above 400 °C. The chemical and thermal stability properties of Ta₂O₅ has made it known as superior materials for sensing properties which lead to high selectivity and sensitivity.

A study was carried out by Zhang et al. [47], where the implementation of Co₃O₄- Ta₂O₅ heterostructure took place in an ethanol gas sensor. It has been seen that this formation possessed an outstanding sensitivity, selectivity, and stability performance which exhibited ethanol gas sensing properties enhancement compared to ZnO-Co₃O₄ and Co₃-SnO₂ heterostructure. This is because Ta₂O₅ nanostructures with ultrathin hollow spherical shells hinder the aggregation of Co₃O₄ nanoparticles while acting as the supporting model. It is claimed to possess more active sites and a higher surface area for adsorption of gas molecules. In this case, aside from Ta₂O₅ selectivity properties, more surface-adsorbed oxygen species are available to react with the target gas, which leads to an enhanced response. In addition, an increase in oxygen chemisorption is also mentioned as being caused by a larger surface area.

Photocatalytic activity

Ta₂O₅ nanostructures exhibit quite promising properties to provide wide surface area and magnificent charge transport and separation. Due to its optical properties which exert a high refractive index, Ta₂O₅ has become a good candidate for photocatalytic application [57], [69]. Several studies have been reported regarding the implementation of Ta₂O₅ related to photocatalytic activity.

Nanostructures of Ta₂O₅ have desired properties in the photoelectrochemical water splitting application as reported by Su et al. [16].

The hydrogen production photocatalytic performance of the heterojunctions created from the $\text{Ta}_2\text{O}_5/\text{g-C}_3\text{N}_4$ is massively better than the pristine $\text{g-C}_3\text{N}_4$. The photocatalytic hydrogen evolution efficiency of the 7.5% TO/CN heterojunction was about 4.2 times higher than that of pure $\text{g-C}_3\text{N}_4$. Furthermore, the 7.5% TO/CN sample demonstrated excellent photochemical stability after 20 hours of photocatalytic testing [54]. Besides, $\text{TiO}_2\text{-Ta}_2\text{O}_5$ (2:1 ratio) photo-catalyst mixed oxide prepared by Patil et al. [62] through the hydrothermal method and calcinated at 600 °C had the highest percent degradation of 99.48% in 60 min. A study reported on $\beta\text{-Ta}_2\text{O}_5$ nanodimple arrays had an electronic band gap of 4.5 eV and a bulk plasmon resonance at 21.4 eV formed via anodization was enhanced by 40% of its plasmonic photocatalysis function after being embedded with gold nanoparticles which proved that surface modification can be employed on anodic Ta_2O_5 to increase the efficiency [56]. Table 3 summarizes the performance of Ta_2O_5 in photocatalytic activity based on anodization, hydrothermal, solvothermal, and sol-gel synthesis methods.

pH sensing

The issues raised from pH sensor fabrication are the complexity of the method, expensive materials, require the reference, other than sensor accuracy and response time which affect the reliability of the sensor. Many studies have been conducted to investigate the use of Ta_2O_5 as a material in pH sensors. However, there is still a gap in the literature of the fabrication of pH sensors from anodized Ta_2O_5 . The mechanism of the pH sensor works is conducted by representing the concentration of hydrogen ions in a solution and portraying it in a significant value known as pH [9]. Ta_2O_5 has often being selected as pH sensor fabrication due to its pH-sensitive properties towards H^+ ions [79]. There were many studies on the fabrication of pH sensors using Ta_2O_5 either using it alone or combined with other oxides for this purpose.

Ta_2O_5 based electrolyte-ion sensitive membrane-oxide-semiconductor (EIOS) pH sensor for the acid environment was made up of thermally grown SiO_2 with Ta deposited using radio-frequency (RF) sputtering. This sensor was tested in the 1 to 10 pH range and obtained a Nernstian value of 56.19 mV/pH [80]. Using a similar method of RF sputtering, the Ta_2O_5 membrane was employed to IrO_2 to enhance the H^+ response towards the pH sensor [81]. It maintains the Nernstian value of -59.447 to -59.504 mV/pH for pH ranging between 2 to 13. The previous performance of Ta_2O_5 showed high sensitivity and low drift which has drawn significant attention regarding the implementation of Ta_2O_5 in this application [82].

Table 3: Performance of Ta₂O₅ in photocatalytic activity based on synthesis method of Tantalum

Synthesis method	Materials	Photocatalytic mechanism	Degradation / evolution rate	Photocurrent	Efficiency	Ref
Anodization	Ta plate coated with Al layer, nanoimprinted with Si and mask anodization to form Ta ₂ O ₅ nanorod array	Water splitting under UV irradiation	-	1.5 ma/cm2 at 1 V versus RHE	0.69% at 2.9 ma cm ⁻²	[58]
	Ta film with thickness of 0.1mm converted to Ta ₃ N ₅ nanotube arrays	Illuminated photoelectrodes with AM 1.5G-simulated sunlight at 100 mw cm-2	10 mv/s	5.9 ma/cm2 at 1.23 VRHE 12.9 ma/cm2 at 1.59 VRHE	-	[16]
	Ta foil with thickness of 1 mm	Degradation of the MB under UV light in 3 different electrolytes containing ethylene glycol (Sample A), glycerol (Sample	Sample C exhibits the highest photocatalytic activity for MB degradation due to the ability of	-	-	[23]

		B) and HF + H ₂ SO ₄ + H ₂ O (Sample C)	electrolyte to generate nanotubes on Ta ₂ O film			
	Ta foil with thickness of 0.25 mm loaded with NiO	Photocatalytic reactions for hydrogen production	7.7 ± 0.3 mmol/h/g	-	-	[22]
Hydrothermal	TiO ₂ -Ta ₂ O ₅	Degradation of methylene blue (MB) under UV irradiation.				[62]
Solvothermal	TaCl ₅ powder turns into Ta ₂ O ₅ powder self-doped with Ta ⁴⁺	Evaluated by degradation rate of Rhodamine B dye as organic pollutant model	-	-	3.6 % at 447 nm	[70]
Sol-gel	Ta Ethoxide liquid turns into Ta ₂ O ₅ white crystal	Evaluated by degradation rate of MB dye	-	1.3 µa/cm ²	-	[97]

Performance enhancement of pH sensor requires high sensitivity, fast response, and low power consumption, where the morphology of porous and nanostructured thick film has a high surface-to-volume ratio due to a large number of adsorption [79]. These properties can increase sensor efficiency. A collective study of pH sensor fabrication based on Ta₂O₅ with good Nernstian value indicates favourable medium of interaction with H⁺ ion provided by Ta₂O₅ structure. The performance of anodized Ta₂O₅ in pH sensing is yet unexplored. However, evidence from the fabrication of pH sensors from anodized amorphous TiO nanotubular generates great output [83]. TiO₂ nanotubes length fabricated at 33 nm to 800 nm by anodizing between 10 to 90 minutes generated response time of less than 30 s and 54.5 mV/pH Nernst response. Hence, the anodization method has the potential to enhance Ta₂O₅ performance towards H⁺ detection.

In pH sensing application, important parameters such as Nernstian response, response time, hysteresis width, drift, offset, and isothermal current need to be highlighted in the resulting yield. These parameters are used to measure the efficiency of the metal oxide pH sensor. Meanwhile, the synthesis method gives a significant influence on electron transfer and ion exchange performance as the outcomes.

When two electrodes are immersed in a test solution, the potential difference between the electrodes is generated. The value is expressed by the Nernst equation where the potential difference is proportional to the pH of the solution. Nernstian potential dependence on solution pH is determined by H⁺/OH⁻ ions. The time required for a sensor to reach more than 90% of its open circuit potential value in equilibrium is defined by the response time. Obtaining different output voltage for a pH electrode measured in the same buffer solution is where the Hysteresis phenomenon occurs. Hence, it is determined by measuring the potential change of a solution after dropping 0.1M HCl or KOH into it [79]. Besides, by calculating the initial and the final voltage difference at a particular pH, drift potential can be determined [84].

The use of Ta₂O₅ as material in pH sensor and its performance are recorded in Table 4 by gathering Nernstian response, response time, hysteresis width, and drift for each pH sensor fabrication [78], [79], [80], [81], [83], [84]. An example of pH sensor fabrication is illustrated in Figure 11a by implementing Ta₂O₅ as a sensing electrode. Figure 11b visualizes the sensing performance of RuO₂-Ta₂O₅ (70:30%) thick film [79].

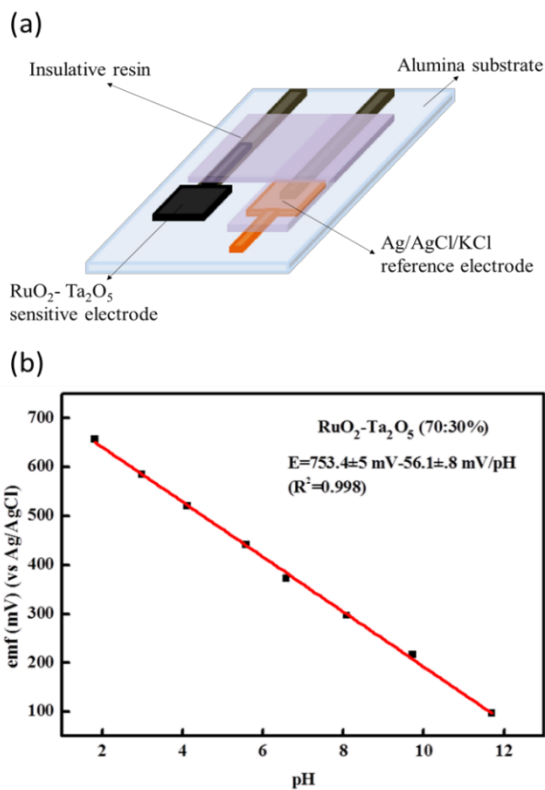


Figure 11. (a) The schematic diagram of pH sensor fabrication using Ta₂O₅ as electrode and (b) RuO₂-Ta₂O₅ (70:30%) thick film sensor potential performance. Reprinted from [79] Copyright 2016, with permission from Elsevier

Table 4: Synthesis method and sensing performance of Ta₂O₅ based pH sensors material

Materials	Synthesis Method	Nernstian response (mV/pH)	Response time (s)	Hysteresis Width	Drift	Ref
Ta ₂ O ₅	Atomic Layer Deposition	53.6 for 39 cycles 55.3 for 200 cycles	-	-	0.017 mV/s	[84]
Ta ₂ O ₅	Thermal oxidation	56.19	-	~5 mV for pH loop 7-4-1-4-7-	-	[80]
IrO ₂ -Ta ₂ O ₅ films on Ta ₂ O ₅	Wet-oxidation	59.447 to 59.504	-	-	-	[81]
Ta ₂ O ₅	Ball milling	45.92	-	-	-	[82]
RuO ₂ - Ta ₂ O ₅	Ball milling	56	>8 s in acidic solutions, >15s in basic solutions	±3 mV for pH loop 2-4-6-4-2, ±8 mV for pH loop 7-9-11-9-7	-	[79]
WASFET with Ta ₂ O ₅ as the top gate oxide layer	Oxidation	58.8 ± 0.2	-	-	< 0.02 mV/h	[85]
Anodized Ta ₂ O ₅ nanotubular	Anodization	31.616 and 24.591	20 s and 90 s	-	-	[98]

Biomedical materials

For the past five years, many kinds of research have been conducted to validate the efficiency of Ta₂O₅ as a sensing layer in biomedical applications. Either being used alone or doped with other materials, Ta₂O₅ is still capable to provide an outstanding outcome during the performance testing. Looking from the perspective of Ta₂O₅ potential as biomedical materials, regardless of the synthesis method, Ta₂O₅ can perform in many various aspects. The fiber-optic sensor has been constructed with a surface plasmon resonance optical scaffold by varying the composition of Ta₂O₅ and Graphene as thin layers with Ta₂O₅ thickness that varies from 2 nm to 12 nm [71]. Sensing probes are prepared with Ta₂O₅ layer, Graphene layer, and both Ta₂O₅-Graphene layers. The combination of Ta₂O₅-Graphene layers resulted in a higher magnitude of field intensity by probes with the Ta₂O₅ layer. This outcome proves that Ta₂O₅ has the ability to enhance fiber-optic sensing performance due to higher sensitivity. Besides, the formation of a biosensor such as a xanthine sensor can be constructed by synthesizing Ta₂O₅ nanostructures through the electrospinning technique [46]. Optimization of a sensing probe made up of xanthine oxide entrapped with Ta₂O₅ nanofibers coating produces a linear operating range of 0 to 3 μM which increases the sensing performance.

Many factors can contribute to its ability to generate reliable outcomes. Besides Ta₂O₅ itself, the anodization method has gained much trust as one of the excellent synthesis methods to enhance Ta₂O₅ properties. Anodization is able to form nanostructures that have the ability to improve in the application by varying its parameter such as electrolyte, voltage applied time interval, and annealing. Other than the evidence of anodized Ti used as coating and bone-implant [20], anodized Ta₂O₅ nanostructures possess outstanding nanomechanical properties to be employed as a coating. Sarraf et al. [32], demonstrated excellent performance of Ta₂O₅ nanotubes with 45 nm inner diameter having the highest corrosion protection of 99% efficiency.

Fabrication of Ta₂O₅ nanoporous through various techniques is capable of imitating the biological properties of human structures which makes it relevant to be implemented in various biomedical applications. The comparison of Ta₂O₅ nanoporous through chemical vapour infiltration and deposition (CVI/CVP) (middle) with powder metallurgy (PM) has physically and mechanically resembled actual human trabecular bone as shown in Figure 9 [14]. Nano-dimpled Ta₂O₅ fabricated via anodization adheres to mouse cranial anterior osteocytes (MC-3T3-E1) cell lines to obtain fluorescence microscopic observation and physicochemical characteristics. Variation in anodization voltage potential increases nano-dimpled structure formation. In contrast to pure Tantalum, increment in pore sizes that can grow between 40 to 90 nm is able to promote the adhesion, proliferation, and differentiation of MC-3T3-E1 cells [12]. This proves that anodization can be used as a synthesis method to enhance Ta₂O₅ performance and obtain the optimum dimension for functional use.

Humidity detection

Humidity detection happens when metal oxide surfaces come into contact with water vapour in the atmosphere or a pure gas [85], [86], [87]. Water adsorption is facilitated in this case by the simultaneous action of the metal oxide's active sites and water molecules. Also, the large surface area of adsorption allows humidity detection to be performed more effectively [89]. In the production of humidity sensors, Ta₂O₅ serves as a sensing layer material. Active sites for interaction media are established on thin-film oxide layers. The surface area available for interaction is influenced by voltage bias and impurities doping. The concentration of water molecules is associated with the moisture level in an atmosphere. In sensor detection, water molecules interact with the oxide layer by adsorption and desorption.

Grotthuss's mechanism principle is widely utilized for most current humidity sensors. In this mechanism, hydrogen bonding serves as the means of proton transfer from one water molecule to its adjacent molecules, resulting in the proton hopping condition [90]-[91]. Surface water at low relative humidity has a high charge carrier density, which attracts hydroxide ions on the surface [91]. The chemisorbed layer is formed due to the adsorption of the hydroxyl group to the Ta₂O₅ oxide layer. An increase in RH increases the water molecules' ability to dissociate, which then releases more hydroxyl groups. Two hydroxyl groups of a water molecule interact with two nearby water molecules by hydrogen bonding and the physisorption (first layer) is formed. When RH concentration increases, the second layer of physisorption is formed on the Ta₂O₅ surface with more water molecules ionized to form H₃O⁺. While the proton transfer has occurred rapidly and H₃O⁺ can separate into H₂O and H⁺, more water adsorption can begin with a further rise in the pH level [10].

For tantalum-based humidity sensor fabricated to this date, Potassium carbonate (K₂CO₃, 99%), niobium pentoxide (Nb₂O₅, 99.5%), tantalum pentoxide (Ta₂O₅, 99%), and manganese dioxide (MnO₂, 99%) was synthesized through a conventional solid-state reaction route and then used to fabricate a tantalum-based resistive humidity sensor, which exhibited a maximum sensitivity of 4600 tested with relative humidity between 15% to 95% at constant 5V bias voltage [86]. The role of tantalum in humidity sensing was once again investigated using the fabrication of tantalum-based interdigital electrodes coated with rGO-SnO₂. By varying 11% to 95% relative humidity over 30 days, this sensor exhibited good sensitivity, rapid response, and long-term stability with the response and recovery times of 10 s and 60 s [92].

More evidence was collected from the nanostructures-based humidity sensor. The sensitivity of the nanoporous alumina anodized in oxalic acid solution has a pore diameter of 34 ± 1 nm and porosity 10 ± 0.5 % that reaches up to 0.21 (pF/%RH). Pore diameter and interpore distance increase the anodic alumina porosity, allowing more water molecules to penetrate and providing more surface area for detection. Thus, the anodic alumina humidity sensor

features improved performance because of its tiny pore structure and uniform nanopore distribution [7]. 40% to 90% relative humidity (RH) was put in contact with a metal-semiconductor-metal humidity sensor made from anodized nanoporous Nb₂O₅ with platinum electrodes, and different bias voltages were applied. The relative sensitivity of 4- μ m-thick Nb₂O₅ nanoporous Nb₂O₅ structure under a voltage bias of 5 V was determined to be 216.5, resulting in the nanoporous Nb₂O₅ presenting the highest sensitivity. This high surface area to volume ratio in the sensors provides highly effective interactions and bindings between active sites and water molecules [3]. This allows the potential of anodized Ta₂O₅ to be further explored as Ta₂O₅ has a good record for H⁺ interaction while considering the nanostructure potential to enhance the water molecules interaction in humidity sensors [99].

Conclusions

In conclusion, the advantage of anodization synthesis method is widely known as it is capable to alter properties according to application requirements by varying its parameter. Ta₂O₅ exhibits outstanding properties as well as thermal and chemical stability which makes it among the best candidate to be implemented in various applications. Anodization of Ta₂O₅ requires a strong etching agent to increase the dissolution rate during pore formation due to its ability to withstand chemical attacks. Hence, this is why Ta₂O₅ is commonly chosen as coating materials aside from being implemented in sensing applications. Nanotubes generation can be enhanced in terms of its length by increasing anodization voltage potential and undergoing thermal treatment afterward. Annealing often takes place after anodization to crystalline as-anodized amorphous state which makes Ta₂O₅ possess extensive properties before fabrication of sensor can be done.

This review justifies the potential of anodization to modify Ta₂O₅ nanostructures and enhance its properties as this method can generate a wide range of nanostructure dimensions that are favourable as medium of interaction. Since Ta₂O₅ has its selectivity to interact with certain molecules, anodization prepares the Ta₂O₅ nanostructures to improve the adhesion and facilitate the sensing process. Various applications have been found by employing Ta₂O₅ either to use it alone or dope with other materials through different synthesis methods such as solvothermal, sol-gel, and ball milling. Future studies will carry out to use Ta₂O₅ as a sensing application as well by performing anodization as a synthesis method to recognize the potential of this method towards Tantalum film. Performance comparison can be made in terms of selectivity, sensitivity, and efficiency for single use of Ta₂O₅ and dope with other materials. By synthesizing the nanostructures through anodization, it is believed that this method can enhance the sensing performance and prove Ta₂O₅ as a reliable metal oxide material for sensing application.

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